

JUMPSTART 4-E

- 1) What charge does Ca like to make?
- 2) What charge does Al like to make?
- 3) What charge does S like to make?
- 4) What charge does Nitrate like to make?
(*hint* you have something in your notebook to help you!)

**Set up p. 99 and p. 101 as KCQ notes!
Targets on the board**

CHEMICAL FORMULAS AND NAMING IONIC COMPOUNDS



Chemical Formulas

- A **chemical formula** is a shorthand way of telling you
 - The **name** of a compound
 - What **type** of **atoms** are in the compound
 - **How many** of each element there are

How to read a formula

- A chemical formula uses
 - **symbols** for each element
 - **subscripts** to tell you how many of each element there are.

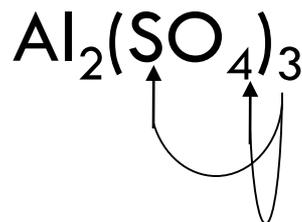
Example:

- If there is no subscript, you assume there is a “1” as the subscript (but you just don’t write it in). N_1H_4

What's with the parentheses?

- If a chemical formula has **parentheses** in it then you have to remember to **distribute** the subscript to each element inside the parentheses.

Example:



There are Aluminum atoms

Sulfur atoms

Oxygen atoms

Practice Questions



Ca = Br =



Cu = N = O =

Naming Ionic Compounds

□ Two types of ionic compounds

□ **BINARY**

- Only TWO types of elements



□ **POLYATOMIC**

- MORE than two types of elements



Naming Binary Compounds

- Cation first, Anion Second
- Metal first, Non-metal Second
- IGNORE THE SUBSCRIPTS!
- Transition metals with more than one possible charge put the charge in parentheses with roman numerals Mn(IV)

- Cation – same name as on periodic table
- Anion – drop the ending and add -ide

NaCl

Practice Naming Binary Compounds

□ AgCl

□ MgO

□ KS

Naming Polyatomic Ionic Compounds

- Cation First, Anion Second
- Both Cation and Anion keep their normal name
- $\text{Mg}(\text{OH})_2$
- $(\text{NH}_4)(\text{NO}_3)$

Practice Naming Ionic Compounds

