**Mixed Ionic Practice**  Name: Period:

Reminders: You need to decide how to name each of these. You know all the rules, but these examples are not sorted.

* Type I means the metal makes a predictable charge (it is colored on your periodic table), no Roman numeral needed.
* Type II means the metal DOES NOT make a predictable charge, you need to figure it out and put a Roman numeral in the name. My suggestion is to circle transition metals to remind yourself.
* Polyatomics are in Table F in your data book. Use the names exactly as they are written. Put parentheses around the polyatomic if you need more than one of them. My suggestion is to underline polyatomic ions to remind yourself.
* Polyatomics can be paired with Type I or Type II metals.
1. NH4Cl
2. Fe(NO3)3
3. TiBr3
4. Cu3P
5. SnSe2
6. Pb(SO4)2
7. Mn2(SO3)3
8. Al(CN)3
9. Na2CO3
10. NaOH
11. MgBr2
12. KCl
13. FeCl2
14. Zn(OH)2
15. BeSO4
16. CrF2
17. Al2S3
18. PbO
19. Li3PO4
20. TiI4
21. Co3N2
22. Mg3P2
23. Ga(NO2)3
24. Ag2SO3
25. NH4OH
26. Al(CN)3
27. Be(CH3COO)2
28. Chromium (VI) phosphate
29. Vanadium (IV) carbonate
30. Tin (II) nitrite
31. Titanium (II) acetate
32. Vanadium (V) sulfide
33. Chromium (III) hydroxide
34. Lithium iodide
35. Lead (II) nitride
36. Silver bromide
37. Sodium phosphide
38. Magnesium nitrate
39. Lead (II) sulfite
40. Calcium phosphate
41. Ammonium sulfate
42. Beryllium chloride
43. Copper (I) arsenide
44. Iron (III) oxide
45. Gallium nitride
46. Iron (II) bromide
47. Vanadium (V) phosphate
48. Calcium oxide
49. Magnesium acetate
50. Aluminum sulfate
51. Copper (I) carbonate
52. Ammonium sulfite