**Organization of the Periodic Table**

1. Elements may react to form ions that have electron configurations like those of the .
2. Which element is in group 15 and period 2? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Which elements are halogens? What charge of ions will they make? Why (what is happening)?
4. Which elements are alkaline earth metals? What charge of ions will they make? Why (what is happening)?
5. List at least three elements that are metals. List three characteristics of metals.
6. List at least three elements that are nonmetals. List three characteristics of nonmetals.
7. Which elements are the metalloids?
8. Describe electronegativity.
9. What happens to the atomic radius as you move left to right across one period? Why?
10. Which element has the highest electronegativity on the whole table?
11. An atom is chemically when all of the orbitals in the outermost energy level are filled.
12. Which group of elements has NO electronegativity and VERY HIGH ionization energy? Why?
13. Is an oxygen ion (oxide) bigger or smaller than an oxygen atom? Why?
14. Is a potassium ion bigger or smaller than a potassium atom? Why?
15. Label the regions of the periodic table with group and period numbers as well as group names. Also label the s, p, d, and f orbital blocks.

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